#### **Pittsburg State University**

#### Pittsburg State University Digital Commons

**Posters** 

2019 Research Colloquium

4-1-2019

### Nanosheets of CuCo<sub>2</sub>O<sub>4</sub> as a High-Performance Electrocatalyst in **Urea Oxidation**

Camila Zequine Pittsburg State University

Ram K. Gupta Pittsburg State University

Pawan K. Kahol Pittsburg State University

Follow this and additional works at: https://digitalcommons.pittstate.edu/posters\_2019



Part of the Energy Systems Commons

#### **Recommended Citation**

Zequine, Camila; Gupta, Ram K.; and Kahol, Pawan K., "Nanosheets of CuCo2O4 as a High-Performance Electrocatalyst in Urea Oxidation" (2019). Posters. 7.

https://digitalcommons.pittstate.edu/posters\_2019/7

This Article is brought to you for free and open access by the 2019 Research Colloquium at Pittsburg State University Digital Commons. It has been accepted for inclusion in Posters by an authorized administrator of Pittsburg State University Digital Commons. For more information, please contact digitalcommons@pittstate.edu.



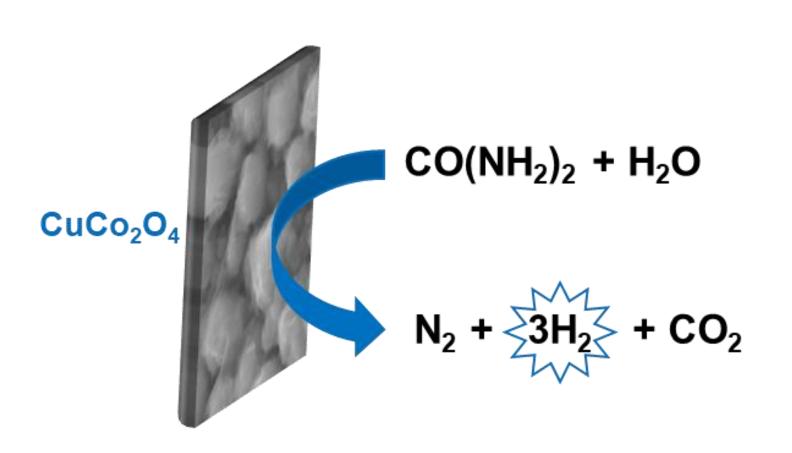
# Nanosheets of CuCo<sub>2</sub>O<sub>4</sub> as a High-Performance Electrocatalyst in Urea Oxidation



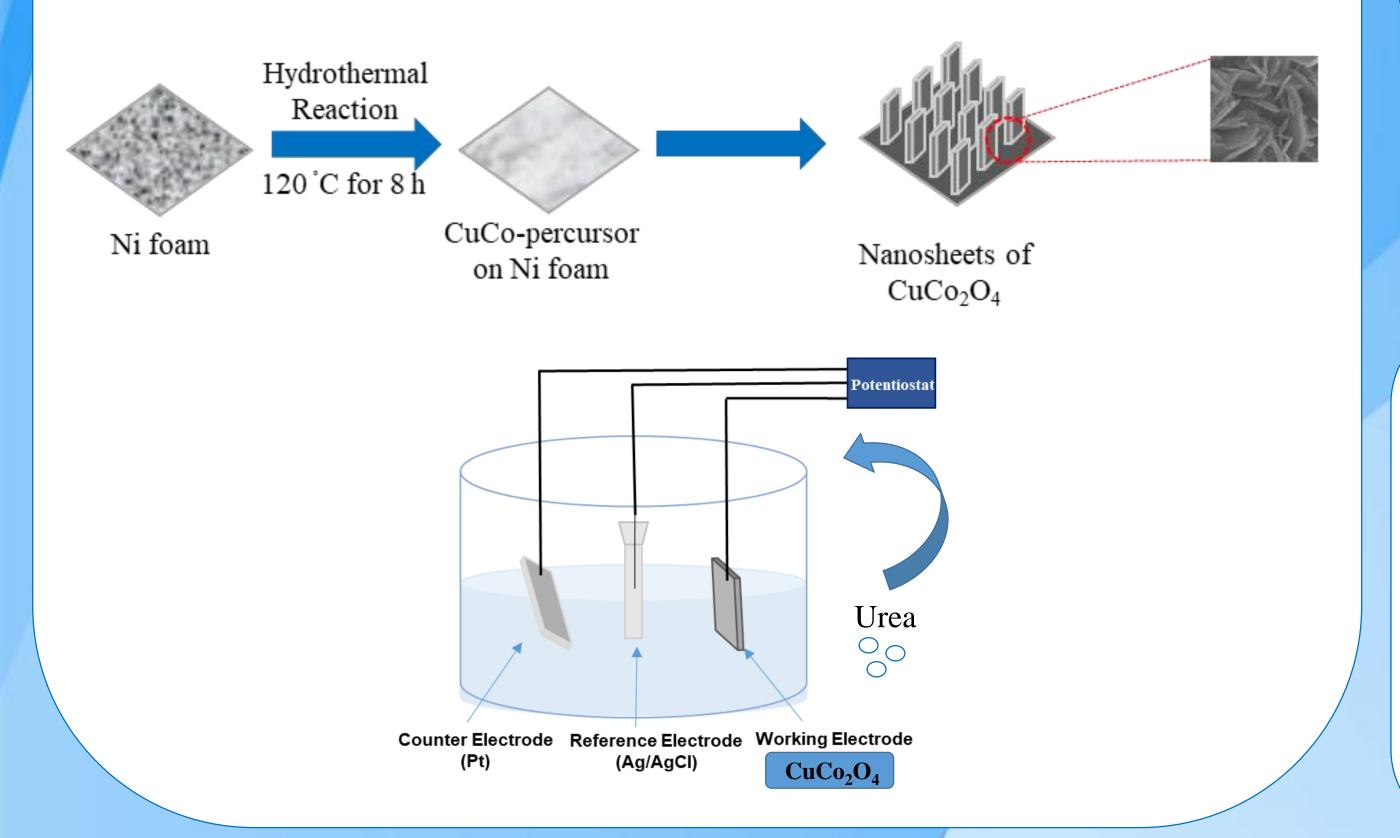
Camila Zequine, Pawan K. Kahol, Ram K. Gupta Pittsburg State University, Pittsburg, KS, United States

### Introduction

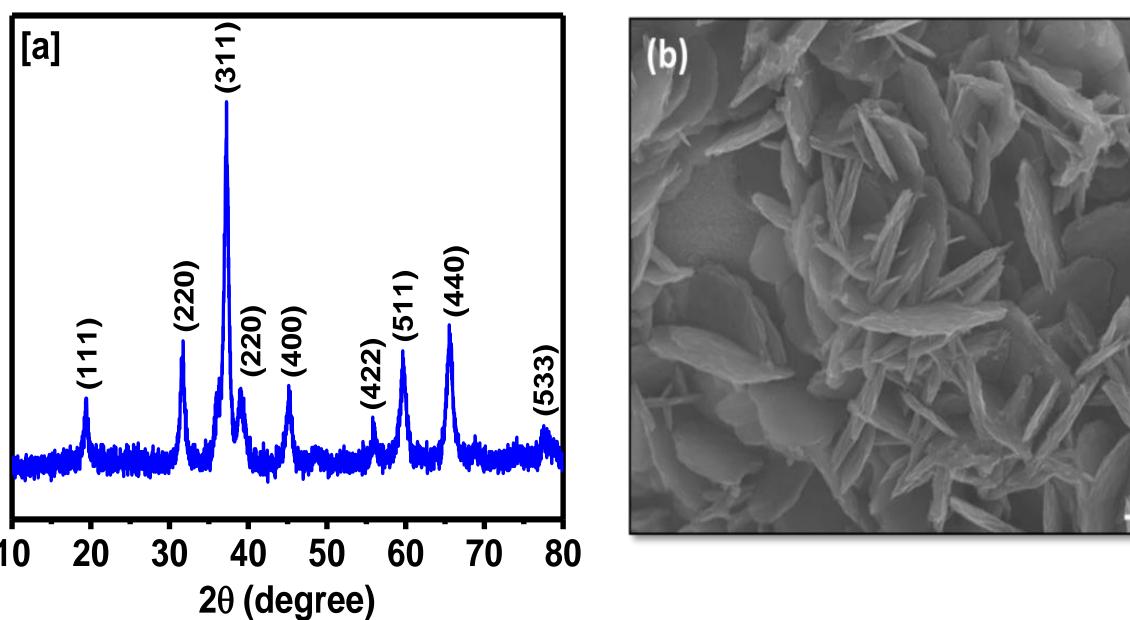
- > Development of new materials capable of generating green energy with high efficiency and low cost.
- abundant waste generated agricultural land which could be used as a source to generate hydrogen.
- > Water splitting occurs at 1.23 V to generate hydrogen, while urea oxidation could provide hydrogen at a lower potential of 0.37 V.
- In this work, it was synthesized nanostructured CuCo<sub>2</sub>O<sub>4</sub> as an electrocatalyst for urea oxidation to generate hydrogen as a green fuel.



# Experimental

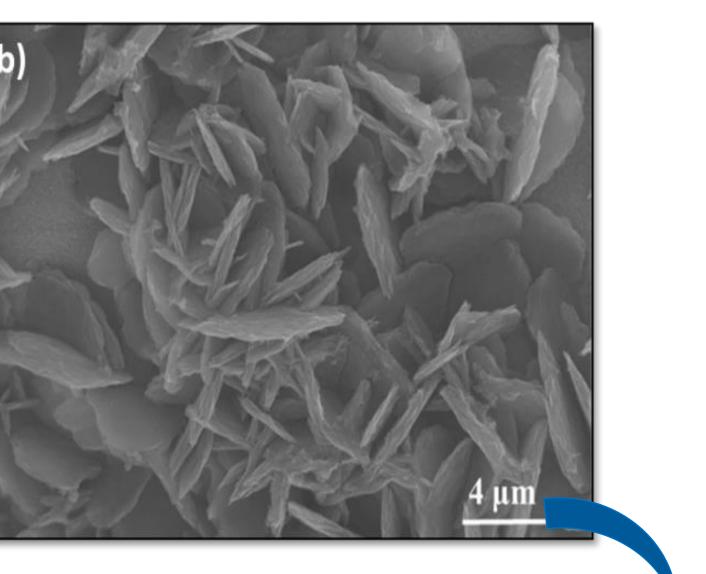


### Results and discussion



(a) XRD pattern and (b) SEM image of the CuCo<sub>2</sub>O<sub>4</sub>

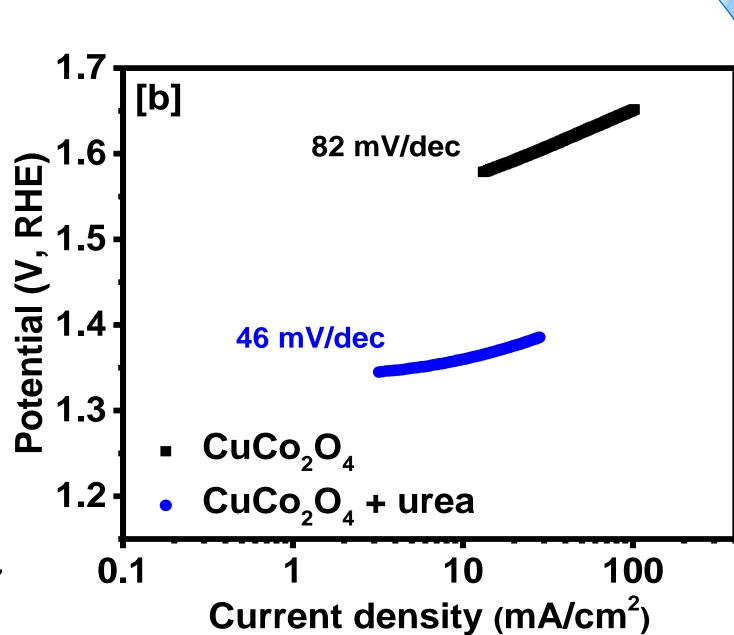
Structure composed of nanosheets in different orientations and a smooth surface.



Potential (V, RHE)

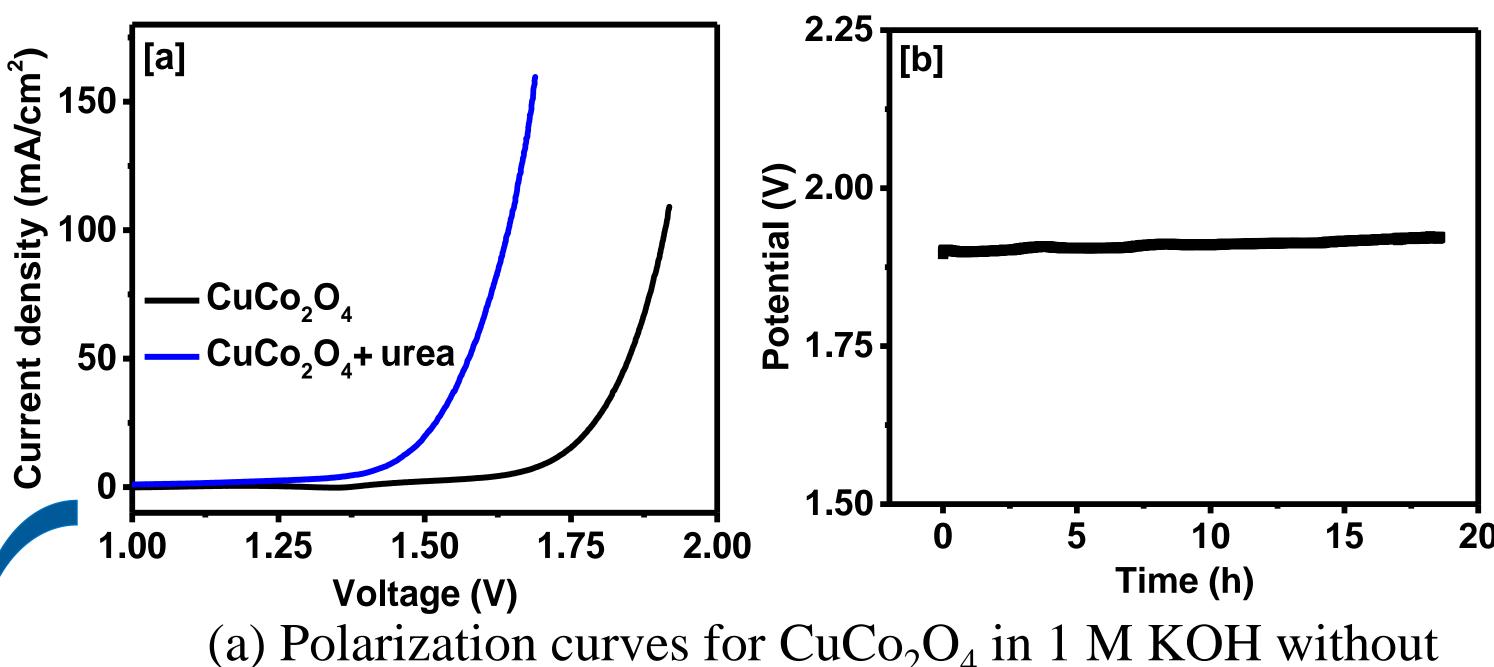
— CuCo<sub>2</sub>O<sub>4</sub>

 $\stackrel{5}{\approx}$  200 \rightarrow CuCo<sub>2</sub>O<sub>4</sub> + urea /



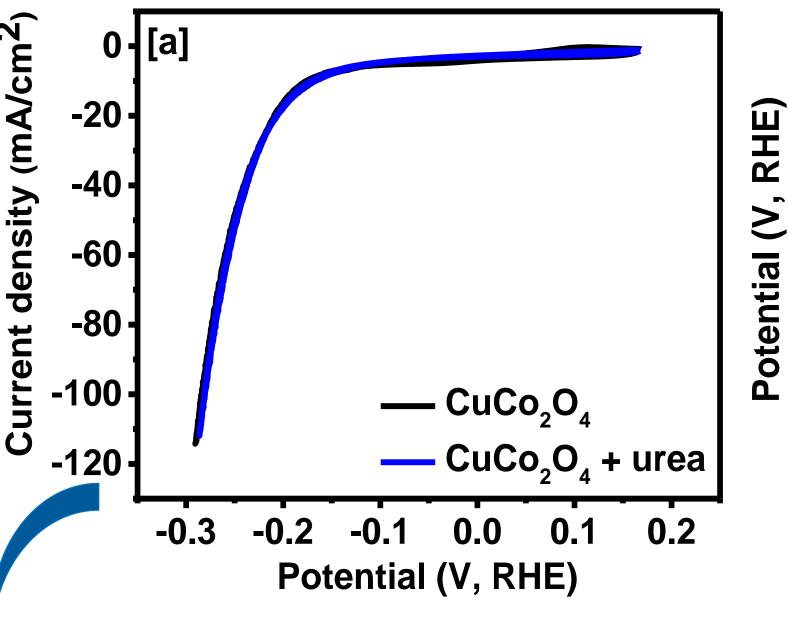
(a) Cyclic voltammetry curves in 1 M KOH with and without 0.33 M urea, and (b) corresponding Tafel slopes.

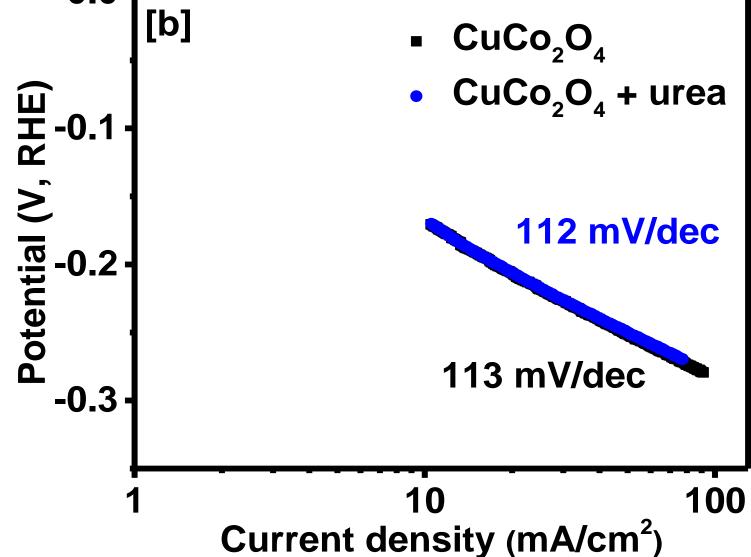
CuCo<sub>2</sub>O<sub>4</sub> demanded a potential of 1.36 V, which is much lower than that without urea.



and with 0.33 M urea, (b) Chronopotentiometric curve for CuCo<sub>2</sub>O<sub>4</sub>.

Electrolyzer requires a cell voltage of 1.45 V in urea solution which is 260 mV less than the cell voltage required without urea.





(a) LSV curves of the CuCo<sub>2</sub>O<sub>4</sub> electrode toward HER in 1.0 M KOH with 0.33 M urea, (b) corresponding Tafel plots.

CuCo<sub>2</sub>O<sub>4</sub> required a potential of 168 mV to generate 10 mA/cm<sup>2</sup>.

### Conclusion and Future work

CuCo<sub>2</sub>O<sub>4</sub> in urea oxidation: a solution to eliminate the urea from groundwater systems and environment that are harmful for people; the H<sub>2</sub> production with less energy consumption and low-costs, which helps to solve the crisis energy.

Future work: Make the process even more efficient and develop new multifunctional materials for energy storage and generation.

# Acknowledgement

Thanks to Polymer Chemistry Program and Kansas Polymer Research for providing research facilities.